

Sample Question Paper 2021-22

Term 1

Subject: Chemistry (043)

Time: 90 Minutes

Max. Marks: 35

General Instructions:

1. The Question Paper contains three sections.
2. Section A has 25 questions. Attempt any 20 questions.
3. Section B has 24 questions. Attempt any 20 questions.
4. Section C has 6 questions. Attempt any 5 questions.
5. All questions carry equal marks. 6. There is no negative marking.

SECTION A

This section consists of 25 multiple choice questions with overall choice to attempt any 20 questions. In case more than desirable number of questions are attempted, ONLY first 20 will be considered for evaluation.

1. Which of the following statements is true:

- (a) Melting point of Phosphorous is less than that of Nitrogen
- (b) N_2 is highly reactive while P_4 is inert
- (c) Nitrogen shows higher tendency of catenation than P
- (d) N-N is weaker than P-P

2. Which of the following is a non-stoichiometric defect?

- (a) Frenkel defect
- (b) Schottky defect
- (c) Metal deficiency defect
- (d) Interstitial defect

3. Identify the law which is stated as:

“For any solution, the partial vapour pressure of each volatile component in the solution is directly proportional to its mole fraction.”

- (a) Henry's law
- (b) Raoult's law
- (c) Dalton's law
- (d) Gay-Lussac's Law

4. Pink colour of LiCl crystals is due to:

- (a) Schottky defect
- (b) Frenkel defect
- (c) Metal excess defect
- (d) Metal deficiency defect

5. Which of the following isomer has the highest melting point:

- (a) 1,2-dichlorobenzene
- (b) 1,3 -dichlorobenzene
- (c) 1,4-dichlorobenzene
- (d) All isomers have same melting points

6. Which one of the following reactions is not explained by the open chain Structure of glucose:

- (a) Formation of pentaacetate of glucose with acetic anhydride.
- (b) Formation of addition product with 2,4 DNP reagent
- (c) Silver mirror formation with Tollen's reagent
- (d) Existence of alpha and beta forms of glucose.

7. Williamson's synthesis of preparing dimethyl ether is an:

- (a) S_N1 reaction
- (b) Elimination reaction
- (c) S_N2 reaction
- (d) Nucleophilic addition reaction

8. Chlorine water loses its yellow colour on standing because:

- (a) HCl gas is produced, due to the action of sunlight.
- (b) A mixture of HOCl and HCl is produced in the presence of light
- (c) HOCl and hydrogen gas is produced
- (d) A mixture of HCl and ClO_3 is produced, due to the action of sunlight

9. During dehydration of alcohols to alkenes by heating with concentrated H_2SO_4 , the initiation step is:

- (a) Protonation of alcohol molecule
- (b) Formation of carbocation
- (c) Elimination of water
- (d) Formation of an ester

10. Amorphous solids are:

- (a) Isotropic
- (b) Anisotropic

- (c) Isotopic
- (d) Isomeric

11. Which of the following reactions is used to prepare salicylaldehyde?

- (a) Kolbe's reaction
- (b) Etard reaction
- (c) Reimer - Tiemann reaction
- (d) Stephen's reduction.

12. Which of the following is an example of a solid solution?

- (a) Sea water
- (b) Sugar solution
- (c) Smoke
- (d) 22 carat gold

13. The boiling points of alcohols are higher than those of hydrocarbons of comparable masses due to:

- (a) Hydrogen bonding
- (b) Ion – dipole interaction
- (c) Dipole- dipole interaction
- (d) Van der Waal's forces.

14. Which of the following has the lowest boiling point:

- (a) H₂O
- (b) H₂S
- (c) H₂Se
- (d) H₂Te

15. Which of the following statement is correct:

- (a) Fibrous proteins are generally soluble in water
- (b) Albumin is an example of fibrous proteins
- (c) In fibrous proteins, the structure is stabilized by hydrogen bonds and disulphide bonds
- (d) pH does not affect the primary structure of protein.

16. Major product obtained on reaction of 3-Phenyl propene with HBr in presence of organic peroxide

- (a) 3- Phenyl 1- bromopropane
- (b) 1 –Phenyl -3- bromopropane
- (c) 1-Phenyl -2-bromopropane
- (d) 3-Phenyl -2- bromopropane

17. Which of the following is a correct statement for C₂H₅Br?

- (a) It reacts with metallic Na to give ethane.
- (b) It gives nitroethane on heating with aqueous solution of AgNO_2
- (c) It gives $\text{C}_2\text{H}_5\text{OH}$ on boiling with alcoholic potash.
- (d) It forms diethylthioether on heating with alcoholic KSH.

18. Covalency of nitrogen is restricted to:

- (a) 2
- (b) 3
- (c) 4
- (d) 5

19. Solubility of gases in liquids decreases with rise in temperature because dissolution is an:

- (a) Endothermic and reversible process
- (b) Exothermic and reversible process
- (c) Endothermic and irreversible process
- (d) Exothermic and irreversible process

20. All elements of Group 15 show allotropy except:

- (a) Nitrogen
- (b) Arsenic
- (c) Antimony
- (d) Bismuth

21. Which of the following is a polysaccharide?

- (a) Glucose
- (b) Maltose
- (c) Glycogen
- (d) Lactose

22. Substance having the lowest boiling point:

- (a) Hydrogen
- (b) Oxygen
- (c) Nitrogen
- (d) Helium

23. Lower molecular mass alcohols are:

- (a) Miscible in limited amount of water
- (b) Miscible in excess of water
- (c) Miscible in water in all proportions

(d) Immiscible in water

24. Maximum oxidation state exhibited by Chlorine is:

(a) +1

(b) +3

(c) +5

(d) +7

25. In which of the following cases blood cells will shrink:

(a) When placed in water containing more than 0.9% (mass/ volume) NaCl solution.

(b) When placed in water containing less than 0.9% (mass /volume) NaCl solution.

(c) When placed in water containing 0.9% (mass/volume) NaCl solution.

(d) When placed in distilled water.

SECTION B

This section consists of 24 multiple choice questions with overall choice to attempt any 20 questions. In case more than desirable number of questions are attempted, ONLY first 20 will be considered for evaluation.

26. How much ethyl alcohol must be added to 1 litre of water so that the solution will freeze at -14°C ? (K_f for water = 1.86°C/mol)

(a) 7.5 mol

(b) 8.5 mol

(c) 9.5 mol

(d) 10.5 mol

27. Which reagents are required for one step conversion of chlorobenzene to toluene?

(a) $\text{CH}_3\text{Cl} / \text{AlCl}_3$

(b) CH_3Cl , Na, Dry ether

(c) $\text{CH}_3\text{Cl}/\text{Fe}$ dark

(d) $\text{NaNO}_2 / \text{HCl} / 0-5^{\circ}\text{C}$

28. On partial hydrolysis, XeF_6 gives:

(a) $\text{XeO}_3 + 4\text{HF}$

(b) $\text{XeO}_2\text{F} + \text{HF}$

(c) $\text{XeOF}_4 + \text{H}_2$

(d) $\text{XeO}_2\text{F}_2 + 4\text{HF}$

29. Which one of the following statements is correct about sucrose:

(a) It can reduce tollen's reagent however cannot reduce fehling's reagent

- (b) It undergoes mutarotation like glucose and fructose
- (c) It undergoes inversion in the configuration on hydrolysis
- (d) It is levorotatory in nature .

30. Phenol does not undergo nucleophilic substitution reaction easily due to:

- (a) Acidic nature of phenol
- (b) Partial double bond character of C- OH bond
- (c) Partial double bond character of C-C bond
- (d) Instability of phenoxide ion

31. Which of the following has highest ionisation enthalpy?

- (a) Nitrogen
- (b) Phosphorus
- (c) Oxygen
- (d) Sulphur

32. Metal M ions form a ccp structure. Oxide ions occupy $\frac{1}{2}$ octahedral and $\frac{1}{2}$ tetrahedral voids. What is the formula of the oxide?

- (a) MO
- (b) MO₂
- (c) MO₃
- (d) M₂O₃

33. The reaction of toluene with Cl₂ in presence of FeCl₃ gives 'X' while the of toluene with Cl₂ in presence of light gives 'Y'. Thus 'X' and 'Y' are:

- (a) X = benzyl chloride, Y = o and p – chlorotoluene
- (b) X = m – chlorotoluene, Y = p – chlorotoluene
- (c) X = o and p–chlorotoluene, Y = trichloromethylbenzene
- (d) X= benzyl chloride, Y = m-chlorotoluene

34. Ozone is a/ an _____ molecule and the two O-O bond lengths in ozone are

(i) _____-and (ii) _____

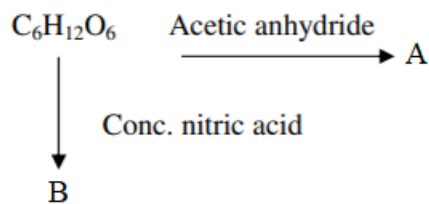
- (a) Linear, 110 pm; 148 pm
- (b) Angular, 110 pm; 148 pm
- (c) Linear, 128 pm; 128 pm
- (d) Angular, 128pm; 128pm

35. Water retention or puffiness due to high salt intake occurs due to:

- (a) Diffusion
- (b) Vapour pressure difference

- (c) Osmosis
 (d) Reverse osmosis

36. In the following reaction, identify A and B:



- (a) A = COOH - (CH₂)₄ - COOH, B = OHC - (CHOCOCH₃)₄ - CH₂OCOCH₃
 (b) A = COOH - (CH₂)₄ - CHO, B = OHC - (CHOCOCH₃)₄ - CH₂OCOCH₃
 (c) A = OHC - (CHOCOCH₃)₃ - CH₂OCOCH₃, B = COOH - (CH₂)₄ - CHO,
 (d) A = OHC - (CHOCOCH₃)₄ - CH₂OCOCH₃, B = COOH - (CH₂)₄ - COOH

37. In lake test for Al³⁺ ions, there is the formation of coloured 'floating lake'. It is due to:

- (a) Absorption of litmus by [Al(OH)₄]⁻
 (b) Absorption of litmus by Al(OH)₃
 (c) Adsorption of litmus by [Al(OH)₄]⁻
 (d) Adsorption of litmus by Al(OH)₃

38. A unit cell of NaCl has 4 formula units. Its edge length is 0.50 nm. Calculate the density if molar mass of NaCl = 58.5 g/mol.

- (a) 1 g/cm³
 (b) 2 g/cm³
 (c) 3 g/cm³
 (d) 4g/cm³

39. Which one of the following are correctly arranged on the basis of the property indicated:

- (a) I₂ < Br₂ < F₂ < Cl₂ (increasing bond dissociation enthalpy)
 (b) H₂O < H₂S < H₂Te < H₂Se (increasing acidic strength)
 (c) NH₃ < N₂O < NH₂OH < N₂O₅ (increasing oxidation state)
 (d) BiH₃ < SbH₃ < AsH₃ < PH₃ < NH₃ (increasing bond angle)

40. What would be the reactant and reagent used to obtain 2, 4-dimethyl pentan-3-ol?

- (a) Propanal and propyl magnesium bromide
 (b) 3-methylbutanal and 2-methyl magnesium
 (c) 2-dimethylpropanone and methyl magnesium iodide

(d) 2- methylpropanal and isopropyl magnesium iodide

41. o-hydroxy benzyl alcohol when reacted with PCl_3 gives the product as (IUPAC name)

(a) o- hydroxy benzyl chloride

(b) 2- chloromethylphenol

(c) o-chloromethylchlorobenzene

(d) 4-hydroxymethylphenol

42. Which of the following statements is true:

(a) Ammonia is the weakest reducing agent and the strongest base among Group 15 hydrides.

(b) Ammonia is the strongest reducing agent as well as the hydrides.

(c) Ammonia is the weakest reducing agent as well as the weakest base among Group 15 hydrides.

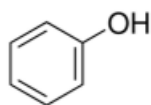
(d) Ammonia is the strongest reducing agent and the weakest base among Group 15 hydrides.

43. Identify the secondary alcohols from the following set:

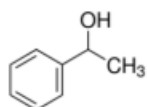
(i) $\text{CH}_3\text{CH}_2\text{CH}(\text{OH})\text{CH}_3$

(ii) $(\text{C}_2\text{H}_5)_3\text{COH}$

(iii)



(iv)



(a) (i) and (iv)

(b) (i) and (iii)

(c) (i) and (ii)

(d) (i), (iii) and (iv)

44. Alkenes decolourise bromine water in presence of CCl_4 due to formation of

(a) Allyl bromide

(b) Vinyl bromide

(c) Bromoform

(d) Vicinal dibromide

45. Given below are two statements labelled as Assertion (A) and Reason (R)

Assertion (A): Electron gain enthalpy of oxygen is less than that of Fluorine but greater than Nitrogen.

Reason (R): Ionisation enthalpies of the elements follow the order Nitrogen > Oxygen > Fluorine

Select the most appropriate answer from the options given below:

- (a) Both A and R are true and R is the correct explanation of A
- (b) Both A and R are true but R is not the correct explanation of A.
- (c) A is true but R is false.
- (d) A is false but R is true.

46. Given below are two statements labelled as Assertion (A) and Reason (R)

Assertion (A): Alkyl halides are insoluble in water.

Reason (R): Alkyl halides have halogen attached to sp^3 hybrid carbon.

Select the most appropriate answer from the options given below:

- (a) Both A and R are true and R is the correct explanation of A
- (b) Both A and R are true but R is not the correct explanation of A.
- (c) A is true but R is false.
- (d) A is false but R is true.

47. Given below are two statements labelled as Assertion (A) and Reason (R)

Assertion (A): Molarity of a solution changes with temperature.

Reason (R): Molarity is a colligative property.

Select the most appropriate answer from the options given below:

- (a) Both A and R are true and R is the correct explanation of A
- (b) Both A and R are true but R is not the correct explanation of A.
- (c) A is true but R is false.
- (d) A is false but R is true.

48. Given below are two statements labelled as Assertion (A) and Reason (R)

Assertion (A): SO_2 is reducing while TeO_2 is an oxidising agent.

Reason (R): Reducing property of dioxide decreases from SO_2 to TeO_2 .

Select the most appropriate answer from the options given below:

- (a) Both A and R are true and R is the correct explanation of A
- (b) Both A and R are true but R is not the correct explanation of A.
- (c) A is true but R is false.
- (d) A is false but R is true.

49. Given below are two statements labelled as Assertion (A) and Reason (R)

Assertion (A): Cryoscopic constant depends on nature of solvent.

Reason (R): Cryoscopic constant is a universal constant.

Select the most appropriate answer from the options given below:

- (a) Both A and R are true and R is the correct explanation of A
- (b) Both A and R are true but R is not the correct explanation of A.
- (c) A is true but R is false.
- (d) A is false but R is true.

SECTION C

This section consists of 6 multiple choice questions with an overall choice to attempt any 5. In case more than desirable number of questions are attempted, ONLY first 5 will be considered for evaluation.

50. Match the following:

I	II
(i)Amino acids	(A)protein
(ii)Thymine	(B)Nucleic acid
(iii)Insulin	(C)DNA
(iv)phosphodiester linkage	(D)Zwitter ion
(v) Uracil	

Which of the following is the best matched options?

- (a) i-A, v- D, iii- C, iv-B
- (b) i-D, ii-C, iii- A, iv-B
- (c) i-D, v- D, iii- A, iv-B
- (d) i-A, ii- C, iii- D, iv-B

51. Which of the following analogies is correct:

- (a) Nitrogen: $1s^2 2s^2 2p^3$:: Argon: $1s^2 2s^2 2p^6$
- (b) Carbon: maximum compounds :: Xenon: no compounds
- (c) XeF_2 : Linear :: ClF_3 : Trigonal planar
- (d) Helium: meteorological observations :: Argon: metallurgical processes

52. Complete the following analogy:

Same molecular formula but different structures: A:: Non superimposable mirror images:

B

- (a) A: Isomers B: Enantiomer
- (b) A: Enantiomers B: Racemic mixture

(c) A: Stereoisomers B: Retention

(d) A: Isomers B: Stereoisomers

CASE 1: Read the passage given below and answer the following questions 53-55

Early crystallographers had trouble solving the structures of inorganic solids using X-ray diffraction because some of the mathematical tools for analyzing the data had not yet been developed. Once a trial structure was proposed, it was relatively easy to calculate the diffraction pattern, but it was difficult to go the other way (from the diffraction pattern to the structure) if nothing was known a priori about the arrangement of atoms in the unit cell. It was important to develop some guidelines for guessing the coordination numbers and bonding geometries of atoms in crystals. The first such rules were proposed by Linus Pauling, who considered how one might pack together oppositely charged spheres of different radii. Pauling proposed from geometric considerations that the quality of the "fit" depended on the radius ratio of the anion and the cation.

If the anion is considered as the packing atom in the crystal, then the smaller cation fills interstitial sites ("holes"). Cations will find arrangements in which they can contact the largest number of anions. If the cation can touch all of its nearest neighbour anions then the fit is good. If the cation is too small for a given site, that coordination number will be unstable and it will prefer a lower coordination structure. The table below gives the ranges of cation/anion radius ratios that give the best fit for a given coordination geometry.

Coordination number	Geometry	$\rho = r_{\text{cation}}/r_{\text{anion}}$
2	linear	0 - 0.155
3	triangular	0.155 - 0.225
4	tetrahedral	0.225 - 0.414
4	square planar	0.414 - 0.732
6	octahedral	0.414 - 0.732
8	cubic	0.732 - 1.0
12	cuboctahedral	1.0

53. The radius of Ag^+ ion is 126pm and of I^- ion is 216pm. The coordination number of Ag^+ ion is:

(a) 2

(b) 3

(c) 6

(d) 8

54. A solid AB has square planar structure. If the radius of cation A^+ is 120pm, calculate the maximum possible value of anion B^-

- (a) 240 pm
- (b) 270 pm
- (c) 280 pm
- (d) 290 pm

55. A “good fit” is considered to be one where the cation can touch:

- (a) All of its nearest neighbour anions.
- (b) Most of its nearest neighbour anions.
- (c) Some of its nearest neighbour anions.
- (d) None of its nearest neighbour anions.