## Assignment No – 1 P – Block Elements

			<u>K Liements</u>	
Que 1. Arrange the followi	ng in order of increasing	g acidity H <sub>3</sub> PO <sub>4</sub> , H <sub>2</sub> CO <sub>3</sub> ,	HCl, HI	
(a) H <sub>3</sub> PO <sub>4</sub> , HCl, H <sub>2</sub> CO <sub>3</sub> , HI	(b) H <sub>3</sub> PO <sub>4</sub> , H <sub>2</sub> CO <sub>3</sub> , HC	Cl, HI (c) $H_2CO_3$ , $H_3F$	PO <sub>4</sub> , HCl, HI (d) None of these	
Que 2. Which of the follow	ing oxides is the most ac	idic?		
(a) N <sub>2</sub> O <sub>5</sub>	(b) P <sub>2</sub> O <sub>5</sub>	(c) $As_2O_5$	(d) Sb <sub>2</sub> O <sub>5</sub>	
Que 3. On heating KClO <sub>3</sub> ,	we get			
(a) $\text{KClO}_2 + \text{O}_2$	(b) KCl + $O_2$	(c) KCl + $O_3$	(d) $KCl + O_2 + O$	
Que 4. Which one is not a	property of ozone?			
(a) it acts an oxidising agent	in dry state	(b) oxi	dation of KI into KIO2	
(c) PbS is oxidised to PbSO <sub>4</sub>	1	(d) Hg	is oxidised to Hg <sub>2</sub> O	
Que 5. The number of P –	O – P bonds in cyclic me	taphosphoric acid is –		
(a) zero	(b) two	(c) three	(d) four	
Que 6. Molecular shapes of	f SF4, CF4 and XeF4 are -	_		
(a) The same with 2, 0 and 1	lone pairs of electrons res	spectively		
(b) The same with 1, 1 and 1	lone pairs of electrons rea	spectively		
(c) Different with 0, 1 and 2	lone pairs of electrons res	pectively		
(d) Different with 1, 0 and 2	lone pairs of electrons res	spectively.		
Que 7. The correct order o	f acidic strength is –			
(a) $Cl_2O_7 > SO_2 > P_4O_{10}$	(b) $CO_2 > N_2O_5 > SO_3$	(c) $Na_2O > MgO > A$	$l_2O_3$ (d) $K_2O > CaO > MgO$	
Que 8. The correct order o	f increasing electron affi	nity of halogens is:		
(a) I < Br < Cl	(b) Br < I < Cl	(c) $Cl < Br < I$	(d) $I < Cl < Br$	
Que 9. Which of the follow	ing is planar?			
(a) XeO <sub>4</sub>	(b) XeO <sub>3</sub> F	(c) $XeO_2F_2$	(d) XeF <sub>4</sub>	
Que 10. Which one of the f	following molecules will f	orm a linear polymeric	structure due to hydrogen bonding?	
(a) HCl	(b) HF	(c) H <sub>2</sub> O	(d) NH <sub>3</sub>	
Que 11. The number of S -	- S bonds in sulphur trioz	xide trimer (S <sub>3</sub> O <sub>9</sub> ) is –		
(a) three	(b) Two	(c) one	(d) zero	
Que 12. Which of the follow	wing statements are corr	ect?		
(a) Among halogens, radius	ratio between iodine and f	luorine is maximum		
(b) Leaving F-F bond, all	halogens have weaker X-	—X bond than X—X' b	ond in interhalogens (c) Among interhalogen	i
compounds maximum numb	per of atoms are present in	iodine fluoride.		
(d) Interhalogen compounds	are more reactive than ha	logen compounds		
Que 13. The number of $\sigma$ -	bonds in P <sub>4</sub> O <sub>10</sub> is –			
(a) 6	(b) 16	(c) 20	(d) 7	
Que 14. In NO <sub>3</sub> <sup>-</sup> ion, the nu	umber of bonds pairs and	d lone pair of electrons	of nitrogen atom are –	
(a) 2, 2	(b) 3, 1 (c) 1, 3	(d) 4, 0	)	
Que 15. Which of the follow	wing has p $\pi$ - d $\pi$ bonding	g?		
(a) NO <sub>3</sub> <sup>-</sup>	(b) $SO_3^{2-}$	(c) $BO_3^{-3}$	(d) $CO_3^{-2}$	
Que 16. Number of lone pa	nirs of electrons on Xe at	oms in XeF2, XeF4 and X	XeF6 molecules are respectively –	
(a) 3, 2 and 1	(b) 4, 3 and 2	(c) 2, 3 and 1	(d) 3, 2 and 0	

Que 17. Which of the following statements are corre	ect for SO <sub>2</sub> gas?		
(a) It acts as bleaching agent in moist conditions	(b) Its mol	ecule has linear geon	netry
(c) Its dilute solution is used as disinfectant. (d) It can	n be prepared by the	e reaction of dilute H	2SO4 with metal sulphide.
Que 18. The true statement for the acids of phospho	rus, H <sub>3</sub> PO <sub>2</sub> , H <sub>3</sub> PO	3 and H3PO4 is –	
(a) The order of their acidity is $H_3PO_4 > H_3PO_3 > H_3PO_4$	(b) All of	them are reducing in	nature
(c) All of them are tribasic acids	(d) The geometry of	of phosphorus is tetral	hedral in all the three
Que 19. Which of the following substances has the h	ighest proton affin	uity?	
	(c) $H_2S$	(d) NH <sub>3</sub>	
Que 20. On boiling phosphorus with KOH solution,	product formed is	_	
(a) Potassium sulphate (b) Phosphorus pentoxide	- e(c) Phosphorus hyd	droxide (d) Ph	nosphine
Que 21. Which of the following are isoelectric and is	ostructural? NO <sub>3</sub> -	, CO <sub>3</sub> <sup>2–</sup> , ClO <sub>3</sub> <sup>–</sup> , SO <sub>3</sub>	-
	(c) $CO_3^{2-}$ , $ClO_3^{-}$		O <sub>3</sub> <sup>2–</sup> , SO <sub>3</sub>
Que 22. Which of the following statements are true?	,		
(a) Only type of interactions between particles of nob	le gases are due to	weak dispersion for	ces. (b) Ionisation enthalpy of
molecular oxygen is very close to that of Xenon.		-	
(c) Hydrolysis of $XeF_6$ is a redox reaction.			
(d) Xenon fluorides are not reactive.			
Que 23. [A] + $H_2SO_4 \rightarrow$ [B] a colourless gas with irr	itating smell [B] +	$K_2Cr_2O_7 + H_2SO_4 -$	$\rightarrow$ green solution. [A] and [B]
are –			<b>u</b>
(a) $SO_3^{2-}$ , $SO_2$ (b) Cl <sup>-</sup> , HCl	(c) $S^{2-}$ , $H_2S$	(d) CO <sub>3</sub> <sup>2–</sup> , CO	2
Que 24. For H <sub>3</sub> PO <sub>3</sub> and H <sub>3</sub> PO <sub>4</sub> , the correct choice is	_		
(a) $H_3PO_3$ is dibasic and reducing	(b) H <sub>3</sub> PO <sub>3</sub> is dibasi	c and non-reducing	
(c) $H_3PO_4$ is tribasic and reducing	(d) H <sub>3</sub> PO <sub>3</sub> is tribas	ic and non-reducing	
Que 25. Total number of lone pairs of electrons in X	eOF <sub>4</sub> is –	C C	
(a) 0 (b) 1	(c) 2	(d) 3	
Que 26. The acid having O–O bond is –			
(a) $H_2S_2O_3$ (b) $H_2S_2O_6$	(c) $H_2S_2O_8$	(d) $H_2S_4O_6$	
Que 27. Among the following, the pair in which the	two species are not	t isostructural is –	
(a) SiF <sub>4</sub> and SF <sub>4</sub> (b) $IO_3^-$ and XeO <sub>3</sub>	(c) BH <sub>4</sub> <sup>-</sup> a	nd NH4 <sup>+</sup>	(d) $PF_6^-$ and $SF_6$
Que 28. In BrF <sub>3</sub> molecule, the lone pairs occupy equ	atorial positions to	) minimise –	
(a) lone pair-bond pair repulsion only	(b) bond p	air- bond pair repulsi	ons only
(c) lone pair-lone pair and lone pair-bond pair repulsior	ıs (d	l) lone pair-lone pair	repulsions only
Que 29. Which of the following is the increasing ord	er of enthalpy of v	aporization?	
(a) NH <sub>3</sub> , PH <sub>3</sub> , AsH <sub>3</sub> (b) AsH <sub>3</sub> PH <sub>3</sub> , NH <sub>3</sub>	(c) NH <sub>3</sub> , A	AsH <sub>3</sub> , PH <sub>3</sub> (d) PH	H <sub>3</sub> , AsH <sub>3</sub> , NH <sub>3</sub>
Que 30. Which of the following oxides of nitrogen is	solid?		
(a) $NO_2$ (b) $H_2O$ (c) $N_2O_2$	3 (d	l) N <sub>2</sub> O <sub>5</sub>	
Que 31. Among Al <sub>2</sub> O <sub>3</sub> , SiO <sub>2</sub> , P <sub>2</sub> O <sub>3</sub> and SO <sub>2</sub> , the corr	ect order of acid st	trength is –	
(a) $SO_2 < P_2O_3 < SiO_2 < As_2O_3$		$SiO_2 < SO_2 < As_2O_2$	$P_3 < P_2O_3$
(c) $As_2O_3 < SiO_2 < SO_2 < P_2O_3$		1) $As_2O_3 < SiO_2 < P_2O_3$	
Que 32. Match the items of column 1 and column 2 a	× ×		
Column 1	Column 2		

(A) Its partial hydrolysis do	es not	(1) He change oxidation state of	central atom
(B) It is used in modern div		(1) The change oxidation state of (2) $XeF_6$	
(C) It is used to provide iner	0 11	(3) XeF <sub>4</sub> for filling electrical but	lbs
<ul><li>(D) Its central atom is in sp<sup>3</sup></li></ul>	-	(4) Ar	103
(a) A–1, B–4, C–2, D–3	d hybridisation	(b) A-1, B-2, C-3, D-4	1
(a) A-1, B-4, C-2, D-3 (c) A-2, B-1, C-4, D-3		(d) $A-1$ , $B-3$ , $C-2$ , $D-4$	
	following footons is most		
- 0	U U	t important in making fluorine t	0 0 0
(a) electron affinity	(b) ionization energy	(c) hydration energy	(d) bond dissociation energy
		ted to exhibit paramagnetic beh	
(a) $CO_2$	(b) ClO <sub>2</sub>	(c) $SO_2$	(d) SiO <sub>2</sub>
		represents the correct order of e	lectron gain enthalpy (with negative
sign) of the given atomic s	-		
(a) $F < Cl < O < S$	(b) $S < O < Cl < F$		(d) Cl < F < S < O
	0	o phosphorus atom in hypophos	•
(a) zero	(b) two	(c) one	(d) three
		orus is thermodynamically most	
(a) Red	(b) White	(c) Black	(d) Yellow
Que 38. Which of the follo	wing contains maximum	number of lone pairs of electron	ns on the central atom?
(a) $\text{ClO}_3^-$	(b) XeF <sub>4</sub>	(c) SF <sub>4</sub>	(d) $I_{3}^{-}$
Que 39. Among the follow	ing molecules		
(i) XeO <sub>3</sub> (ii) Xe	OF <sub>4</sub> (iii) Xe	$eF_6$	
those having same number of	of lone pairs on Xe are –		
(a) (i) and (ii) only	(b) (i) and (iii) only	(c) (ii) and (iii) only	(d) (i), (ii) and (iii)
Que 40. The number of	P – O – P bonds in the	e structure of phosphorus pent	oxide and phosphorus trioxide are
respectively –			
(a) 6, 6	(b) 5, 5	(c) 5, 6	(d) 6, 5
Que 41. There are no S – S	5 bond in –		
(a) $S_2O_4^{2-}$	(b) $S_2O_5^{2-}$	(c) $S_2O_3^{2-}$	(d) $S_2O_7^{2-}$
Que 42. In which of the fol	llowing molecules/ ions a	re all the bonds not equal?	
(a) SF <sub>4</sub>	(b) SiF <sub>4</sub>	(c) XeF <sub>4</sub>	(d) $BF_4^-$
Que 43. Which products a	re expected from the disj	proportionation reaction of hype	ochlorous acid?
(a) HClO <sub>3</sub> and Cl <sub>2</sub> O	(b) HClO <sub>2</sub> and HClO <sub>4</sub>	(c) HCl and Cl <sub>2</sub> O	(d) HCl and HClO <sub>3</sub>
Que 44. The decreasing va	alues of bond angles for I	NH <sub>3</sub> (106) to SbH <sub>3</sub> (101 <sup>0</sup> ) down	group 15 of the periodic table is due
to –			
(a) Increasing bond pair-bon	nd pair repulsion	(b) Increasing p-orbital	character in sp <sup>3</sup>
(c) Decreasing bond pair-bo	ond pair repulsion	(d) Decreasing electron	egativity
Que 45. Which of the follo	wing statements is true?		
(a) H <sub>3</sub> PO <sub>3</sub> is a stronger acid	than H <sub>2</sub> SO <sub>3</sub>	(b) In aqueous medium, HF is a	stronger acid than HCl
(c) HClO <sub>4</sub> is a weaker acid	than HClO <sub>3</sub>	(d) HNO <sub>3</sub> acid is a stronger acid	than HNO <sub>2</sub>
Que 46. Which of the follo	wing chemical reactions	depicts the oxidizing behaviour	of H <sub>2</sub> SO <sub>4</sub> ?
(a) 2 HI + H <sub>2</sub> SO <sub>4</sub> $\rightarrow$ I <sub>2</sub> + SO	$O_2 + 2 H_2O$	(b) Ca $(OH)_2 + H_2SO_4 \rightarrow CaSO$	$_4 + 2 H_2O$

<form>Assertion and Reason I ye Note: In the following questions a statement of assertion followed by a statement of reason is given. Choose the correct answer out of the following cuestions at correct statements and reason is not the correct explanation of the assertion. (a) Both assertion and reason are correct statements, but reason is not the correct explanation of the assertion. (b) Both assertion is or correct statements, but reason is not the correct explanation of the assertion. (c) Assertion is or correct, but reason is correct statements. (d) Assertion is or or exproved by the reaction of KI with concentrated H<sub>3</sub>SO<sub>1</sub>. Que 47. Assertion: II currect by the prepared by the reaction of KI with concentrated H<sub>3</sub>SO<sub>1</sub>. Que 48. Assertion: Both trends atom guitare exist as S<sub>1</sub> but oxygen exist as O<sub>2</sub>. Reason: Hhas lowest H— X both strength atom guitare exist as S<sub>1</sub> but oxygen exist as O<sub>2</sub>. Reason: Coxygen forms µ— µ = µ multiple bond due to Que 49. Nhich is the strength atom guitare exist as S<sub>1</sub> but oxygen exist as O<sub>2</sub>. Reason: Oxygen forms µ— µ = µ multiple bond due to Que 50. Which of the following outside of nitrogen is thermally most statle? (a) N<sub>2</sub>O<sub>2</sub> (b) N<sub>2</sub>O (c) NO (d) N<sub>2</sub>O<sub>3</sub>. Que 51. Which is the strength atomic for Support (d) N<sub>2</sub>O<sub>3</sub>. Que 52. The black ing a correct Support (d) C (d) N<sub>2</sub>O<sub>3</sub>. (a) ascent oxygen (b) cholmine (c) (C) RO (d) O (d) RC (d) C) Que 53. Nitrogen forms ¥=i very of compounds in tates rate form - (a) 1 ppm (b) 3 ppm (c) 5 pom (d) 1 ppm (d) Ppm (d) Ppm (d) 1 ppm (d) PC (d) (d)</form>
answer out of the following choice.(a) Both assertion and reason are correct statements, and reason is the correct explanation of the assertion.(b) Both assertion and reason are correct statements, but reason is not the correct explanation of the assertion.(c) Assertion is correct, but reason is correct statement.(d) Assertion is correct attements.(e) Both assertion and reason are wrong statement.(e) Both assertion: H1 cannot be prepared by the reaction of K1 with concentrated H₂SO₄.Reason: H1 has lowest H—X bond strength among halogen acids.Que 47. Assertion: Both rhombic and monoclinic sulphur exist as S <sub>b</sub> but oxygen exist as O₂.Reason: Oxygen forms pa – pa multiple bond due toQue 49. Which is the stronget acid?(a) H₂SO₄(b) HC1(c) NC0(d) HNO3Que 50. Which of the following oxides of nitrogen is thermally most stable?(a) N₂O₂(b) N₂O(c) NC(d) KrQue 51. Which inert gas has abnormal behaviour on Higuefaction?(a) N₂O₂(b) N₂O(a) N₂O₂(b) Alo?(b) Slop(c) Ar(c) BrtS_0(d) KrQue 52. The bleaching active of CoOL3 is all to iddition states ranging from -(a) 3 to + 5(b) 3 to + 3(c) 3 to + 4(d) 3 to + 6Que 53. The consent with thas molecular nature in gas phase but olici at sold states is -(a) PCI₂(b) Al20(c) S C pipm(d) H2O₂(d) S C pipm(d) D pipmQue 54. High concentration for the stable?(a) Seo₂(b) Al20(c) S pipm(d) D goo]
(a) Both assertion and reason are correct statements, but reason is not the correct explanation of the assertion.   (b) Both assertion is correct, but reason is wrong statement.   (c) Assertion is correct, but reason is correct statements, but reason is not the correct explanation of the assertion.   (c) Assertion is correct, but reason is wrong statement.   (d) Assertion is wrong but reason is correct statements.   (d) Both assertion and reason are wrong statements.   (e) 43. Assertion: Both more be prepared by the reaction of KI with concentrated H <sub>2</sub> SO <sub>4</sub> .   Reason: Both assertion and monoclinic sulptur exist as S <sub>b</sub> but oxygen exist as O <sub>2</sub> .   Reason: Oxygen forms pa – p multiple bond due to   Que 48. Assertion: Both more statements is thermally most stable?   (a) N <sub>2</sub> O <sub>3</sub> (b) HCI (c) HCO <sub>4</sub> (d) HNO <sub>3</sub> Que 51. Which in the stand monoclinic sulptane exist as S <sub>b</sub> but oxygen exist as O <sub>2</sub> . (d) N <sub>2</sub> O <sub>3</sub> (a) N <sub>2</sub> O <sub>3</sub> (b) N <sub>2</sub> O (c) NO (d) HNO <sub>3</sub> Que 50. Which of the following oxides of nitrogen is thermally most stable? (d) N <sub>2</sub> O <sub>3</sub> (a) N <sub>2</sub> O <sub>3</sub> (b) N <sub>2</sub> O (c) NO (d) N <sub>2</sub> O   (a) So D <sub>3</sub> (b) N <sub>2</sub> O (c) Cl CO (d) Cl T   (a) ascent oxygen (b) AlpO <sub>3</sub> (c) AlpO <sub>3</sub> (d) - 3 to A   (a) ascent oxyge
(b) Both assertion and rearrow are correct statements, but reason is not the correct explanation of the assertion.   (c) Assertion is correct, but reason is correct statement.   (d) Assertion is wrong but reason is correct statement.   (e) Both assertion and reason are wrong statements.   Que 51. Assertion: HI has lowest H—x bond strength among halogen acids.   Que 43. Assertion: Both hombic and monoclinic sulput reasts as by but oxygen exist as O <sub>2</sub> .   Reason: Oxygen forms pat – pa multiple bond due to   Que 49. Which is the strengts acid?   (a) N <sub>2</sub> O <sub>3</sub> (b) HCl (c) HCO <sub>4</sub> (d) HNO <sub>3</sub> Que 51. Which instent strengts and the oxides of nitrogen is thermally most stable? (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which instent strengts anong halogen acids. (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which instent strengts anong halogen acids. (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which instent strengts anong halogen acids. (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which instent strengts anong halogen acids. (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which instent strengts anong halogen acids. (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which instent strengts anong halogen acids anong halogen acids. (d) N <sub>2</sub> O <sub>3</sub> (d) N <sub>2</sub> O <sub>3</sub> Que 51. Which of the followent acids ano
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(d) Assertion is wrong but reason is correct statement.   (e) Both assertion and reason is are wrong statements.   Que 47. Assertion: H cambe prepared by the reaction of K1 with concentrated H <sub>2</sub> SO <sub>4</sub> .   Reason: H1 has lowest H—X bond strength among bale gren acids.   Que 48. Assertion: Both there at a monoclinic subture exist as S <sub>8</sub> but oxygen forms pr – pr multiple bond due to:   Que 49. Which is the strength among bale gren acids.   Que 49. Which is the strength among bale gren acids.   Que 49. Which is the strength among bale gren acids.   Que 49. Which is the strength among bale gren acids.   Que 49. Which is the strength among bale gren acids.   Que 49. Which is the strength among bale grent gr
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Que 47. Assertion: HI can-or be prepared by the reaction of KI with concentrated H3SO.Reason: HI has lowest H—X bond strength among bile en acids.Que 48. Assertion: BothW bile and monoclinic with we sist as S <sub>b</sub> but oxygen exist as O <sub>2</sub> .Reason: Oxygen forms pr - pr multiple bond due toQue 49. Which is the strowest acid?Que 49. Which is the strowest acid?Que 49. Which of the bile with concentrated monoclinic with the monoclinic wit
Reason: HI has lowest H—X bond strength among halogen acids.Que 48. Assertion: Both Hombic and monoclinic sub-lur exist as S <sub>8</sub> but oxygen exist as O <sub>2</sub> .Reason: Oxygen forms $p\pi - p\pi$ multiple bond due toQue 49. Which is the strengts acid?(a) H <sub>2</sub> SO <sub>4</sub> (b) HCl(c) HClO <sub>4</sub> (d) HNO <sub>3</sub> Que 50. Which of the following oxides of nitrogen is thermally most stable?(a) N <sub>2</sub> O <sub>5</sub> (b) N <sub>2</sub> O(c) NO(d) N <sub>2</sub> O <sub>3</sub> Que 51. Which inert gas has abnormal behaviour or liguefaction?(a) Xe(b) He(c) Ar(d) KrQue 52. The bleaching actor of CaOCl <sub>2</sub> is due to -(a) naccent oxygen(b) chlorine(c) Ar(d) HClQue 53. Nitrogen forms a variety of compounds in atles rates rate from - to a or 4(a) - 3 to + 5(b) - 3 to + 3(c) - 3 to + 4(d) - 3 to + 6Que 55. The compound har mature in gas phase but ionic state is a sole over -(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound with has molecular nature in gas phase but ionic state is -(a) SeO <sub>2</sub> (b) Al <sub>2</sub> O <sub>3</sub> (c) SO <sub>2</sub> O <sub>3</sub> (d) Bl <sub>2</sub> O <sub>3</sub> Que 55. The compound with state is not in accord and be readel against it?(a) FO <sub>2</sub> (b) Al <sub>2</sub> O <sub>3</sub> (c) SO <sub>2</sub> O <sub>3</sub> (d) Bl <sub>2</sub> O <sub>3</sub> Que 55. The compound with state is not in accord accord against it?(a) SeO <sub>2</sub> (b) Al <sub>2</sub> O <sub>3</sub> (c) SO <sub>2</sub> O <sub>3</sub> </td
Que 48. Assertion: Both - bortholic is unborned line: sub-unit exist as Sa but oxygen exist as O <sub>2</sub> .Reason: Oxygen forms par – par multiple bond due toQue 49. Which is the structed assert of the colloge of the structed assert of the colloge of the
Reason: $Oxygen forms p = -p \pi$ multiple bond due toQue 49. Which is the str=vest stacid?(a) H <sub>3</sub> SO <sub>4</sub> (b) HCl(c) HClO <sub>4</sub> (d) HNO <sub>3</sub> Que 50. Which of the following oxides of nitrogen is thermally most stable?(a) N <sub>2</sub> O <sub>5</sub> (b) N <sub>2</sub> O(c) NO(d) N <sub>2</sub> O <sub>3</sub> Que 51. Which inert gas behaviour us liquefaction?(a) N <sub>2</sub> O(b) He(c) Ar(d) KrQue 52. The bleaching actuation of CaOCl <sub>2</sub> is due to -(d) MCl(d) HClQue 53. Nitrogen forms = writery of compounds is us oxidation states range from -(d) - 3 to + 6(a) - 3 to + 5(b) - 3 to + 3(c) - 3 to + 4(d) - 3 to + 6Que 54. High concentrative fluoride are poisowation states range from -(d) 10 ppm(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound with thas molecular nature is gas hase but ionic in solid state is -(a) SeO <sub>2</sub> (b) Al <sub>2</sub> O <sub>3</sub> (c) PCl <sub>3</sub> (d) POCl <sub>3</sub> Que 57. Which of the following is the most basic(c) PCl <sub>3</sub> (d) Bi <sub>2</sub> O <sub>3</sub> Que 57. Which of the following is the most basic(d) SeO <sub>2</sub> (d) Bi <sub>2</sub> O <sub>3</sub> (a) F2 <sub>2</sub> Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>5</sub> ; bond dissociation energy(b) Sr <sub>2</sub> Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>5</sub> ; oxidizing power(c) H1 > HBr > HCl > HF; ardiic property power(c) BF <sub>3</sub> (d) Al <sub>3</sub> GQue 59. In which of the toil wing mainer, the tow server wing mainer wing
Que 49. Which is the structure statical or state or sta
(a) $H_3SO_4$ (b) $HCl$ (c) $HClO_4$ (d) $HOO_3$ Que 50. Which of the following sides of nitrogen is termally most states(a) $N_2O_5$ (b) $N_2O$ (c) $NO$ (d) $N_2O_3$ Que 51. Which inert gas has aborrmal behaviour ingrefactions?(a) $Xe$ (b) $He$ (c) $Ar$ (d) $Kr$ Que 52. The bleaching is to GAOCl2 is due to(a) nascen oxygen(b) chlorine(c) $HClO$ (d) $HCl$ Que 53. Nitrogen forms a viety of compounds is all visitation states range from(a) $-3$ to $+5$ (b) $-3$ to $+3$ (c) $-3$ to $+4$ (d) $-3$ to $+6$ Que 55. Hie compound with as molecular nature is gas hase but iotic state is(a) $1$ ppm(b) $Sl_2$ (c) $PCl_3$ (d) $POCl_3$ Que 56. Which of the following is the most basic-viete:(a) SeO_2(b) $Al_2O_3$ (c) $Sb_2O_3$ (d) $Bi_2O_3$ Que 57. Which of the following is the most basic viete:(d) $F_2 > Br_2 > I_2$ ; oxidizing power(a) $F_2 > Cl_2 > Br_2 > I_2$ ; bird if write property power(b) $F_2 > Cl_2 > Br_2 > I_2$ ; oxidizing power(c) $HI > HBr > HCI > HF:$ (c) $Br_3$ (c) $Br_3$ (d) $AlF_3$ Que 55. In which of the following molecules, are alised to trace viet subtrace viet viet viet viet viet viet viet vie
Que 50. Which of the following oxides of nitrogen is thermally most stable?(a) $N_2O_5$ (b) $N_2O$ (c) NO(d) $N_2O_3$ Que 51. Which inert gas has abnormal behaviour liquefaction?(a) $Xe$ (b) He(c) Ar(d) KrQue 52. The bleaching action of CaOCl <sub>2</sub> is due to(a) Ac(b) chlorine(c) HClO(d) HClQue 53. Nitrogen forms a variety of compounds in all oxidation states ranging from(a) $-3$ to $+5$ (b) $-3$ to $+3$ (c) $-3$ to $+4$ (d) $-3$ to $+6$ Que 54. High concentration of fluoride are poisonur and harmful to bones and teeth at levels over(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound with has molecular nature in gas phase but ionic in solid state is -(a) PCl_3(d) POCl_3Que 56. Which of the Following is the most basic oxide?(d) POCl_3(d) POCl_3Que 57. Which of the following orders is not in accordance with the property stated against it?(d) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>2</sub> ; oxidizing power(c) HI > Hz > HCl > HF; acidic property power(c) BF <sub>3</sub> (d) AlF <sub>3</sub> (c) HI > Hz > HCl > HF; acidic property power(c) BF <sub>3</sub> (d) AlF <sub>3</sub> (q) NF <sub>3</sub> (b) ClF <sub>3</sub> (c) BF <sub>3</sub> (d) AlF <sub>3</sub> (q) NF <sub>3</sub> (b) ClF <sub>3</sub> (c) So <sup>2</sup> - and NO <sub>3</sub> <sup>-</sup> (d) BF <sub>3</sub> and NF <sub>3</sub> <sup>-</sup>
(a) N2O5(b) N2O(c) NO(d) N2O3Que 51. Which inert gas has abnormal behaviour unteraction?(a) Xe(b) He(c) Ar(d) KrQue 52. The bleaching actor of CaOCl2 is due to(a) nascent oxygen(b) chlorine(c) HClO(d) HClQue 53. Nitrogen forms a variety of compounds in all oxidation states ranging from -(d) -3 to + 6Que 54. High concentration of fluoride are poisonary at the poison of the pois
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Que 52. The bleaching action of CaOCL2 is due to(a) nascent oxygen(b) chlorine(c) HClO(d) HClQue 53. Nitrogen forms a variety of compounds in all oxidation states ranging from -(a) $-3$ to $+5$ (b) $-3$ to $+3$ (c) $-3$ to $+4$ (d) $-3$ to $+6$ Que 54. High concentration of fluoride are poisonary and harmful to bones and teeth at levels over -(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound with has molecular nature in gas phase but ionic is a tate is -(a) PCl5(b) CCl4(c) PCl3(d) POCl3Que 56. Which of the following is the most basic varieties(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the following is not in accurate with the property stated against it?(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the following noters is not in accurate with the property stated against it?(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the following noters is not in accurate with the property stated against it?(a) SeO2(b) Al2O3(c) Sb2O12 > Br2 > I2; boxidizing power(c) H1 > HBr > HCl > HF; acidic property power(d) F2 > Cl2 > Br2 > I2; coxidizing power(c) H1 > HBr > HCl > HF; acidic property power(d) BF3(d) AlF3Que 59. In which of the following mains, the two species are isostructural?(d) AlF3(a) BrO3 and XeO3(b) SF4 and XeF4(c) SO3 <sup>2</sup> and NO3 <sup>-</sup> (d) BF3 and NF3 <sup>-</sup>
(a) nascent oxygen(b) chlorine(c) HClO(d) HClQue 53. Nitrogen forms a variety of compounds in variety in variety of compounds in variety in variety of compounds in variety in variety of compounds in variety of compound variety of fluoride are poisonal and harmful to borse and teeth at levels over –(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound variety has molecular nature in gas phase but ion state is –(a) PCls(b) CCl4(c) PCl3(d) POCl3Que 56. Which of the following is the most basic variety(a) SeO2(b) Al <sub>2</sub> O3(c) Sb <sub>2</sub> O3(d) Bi <sub>2</sub> O3Que 57. Which of the following orders is not in accurate view to the variety stated against it?(a) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>2</sub> ; born d' isoociation energy(b) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>2</sub> ; oxidizing power(c) HI > HBr > HCl > HF; ardic property power(d) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>2</sub> ; electronegativity(a) NF <sub>3</sub> (b) CIF <sub>3</sub> (c) BF <sub>3</sub> (d) AlF <sub>3</sub> Que 59. In which of the following pairs, the two server
Que 53. Nitrogen forms a variety of compounds in all oxidation states ranging from –(a) $-3$ to $+5$ (b) $-3$ to $+3$ (c) $-3$ to $+4$ (d) $-3$ to $+6$ Que 54. High concentration of fluoride are poisonous and harmful to bones and teeth at levels over –(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound with has molecular nature in gas phase but ionic in solid state is –(a) PCl <sub>5</sub> (b) CCl <sub>4</sub> (c) PCl <sub>3</sub> (d) POCl <sub>3</sub> Que 56. Which of the following is the most basic value?(a) SeO <sub>2</sub> (b) Al <sub>2</sub> O <sub>3</sub> (c) Sb <sub>2</sub> O <sub>3</sub> (d) Bi <sub>2</sub> O <sub>3</sub> Que 57. Which of the following orders is not in accordance with the property stated against it?(a) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>2</sub> ; bond dissociation energy(b) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > I <sub>2</sub> ; oxidizing power(c) HI > HBr > HCl > HF; acdic property power(d) BF <sub>3</sub> (d) AlF <sub>3</sub> Que 59. In which of the following pairs, the two species are isostructural?(a) NF <sub>3</sub> (b) CIF <sub>3</sub> (c) SO <sub>3</sub> <sup>2</sup> and NO <sub>3</sub> <sup>-</sup> (d) BF <sub>3</sub> and NF <sub>3</sub> <sup>-</sup>
(a) $-3$ to $+5$ (b) $-3$ to $+3$ (c) $-3$ to $+4$ (d) $-3$ to $+6$ Que 54. High concentration of fluoride are poisonous and harmful to bones and teeth at levels over $-$ (a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound which has molecular nature in gas phase but ionic in solid state is $-$ (a) PCl <sub>5</sub> (b) CCl <sub>4</sub> (c) PCl <sub>3</sub> (d) POCl <sub>3</sub> Que 56. Which of the following is the most basic oxtler:(a) SeO <sub>2</sub> (b) Al <sub>2</sub> O <sub>3</sub> (c) Sb <sub>2</sub> O <sub>3</sub> (d) Bi <sub>2</sub> O <sub>3</sub> Que 57. Which of the following orders is not in accordance with the property stated against it?(a) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > l <sub>2</sub> ; bond dissociation energy(b) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > l <sub>2</sub> ; oxidizing power(c) H1 > HBr > HCl > HF; acidic property power(d) F <sub>2</sub> > Cl <sub>2</sub> > Br <sub>2</sub> > l <sub>2</sub> ; electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) NF <sub>3</sub> (b) CIF <sub>3</sub> (c) BF <sub>3</sub> (d) AlF <sub>3</sub> Que 59. In which of the following pairs, the two species are isostructural?(a) BrO <sub>3</sub> and XeO <sub>3</sub> (b) SF <sub>4</sub> and XeF <sub>4</sub> (c) SO <sub>3</sub> <sup>2-</sup> and NO <sub>3</sub> <sup>-</sup> (d) BF <sub>3</sub> and NF <sub>3</sub> <sup>-</sup>
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(a) 1 ppm(b) 3 ppm(c) 5 ppm(d) 10 ppmQue 55. The compound what has molecular nature in gas phase but ionic(a) PCl5(b) CCl4(c) PCl3(d) POCl3Que 56. Which of the tollow is the most basic(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the follow is orders is not in accertaint with the property stated against it?(a) $F_2 > Cl_2 > Br_2 > I_2;$ boord is sociation energy(b) $F_2 > Cl_2 > Br_2 > I_2;$ oxidizing power(c) HI > HBr > HCl > HF; acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2;$ electronegativityQue 58. In which of the Follow is molecules, are isostructured with the state is order with the stat
Que 55. The compound which has molecular nature in gas phase but ionic in solid state is –(a) PCl5(b) CCl4(c) PCl3(d) POCl3Que 56. Which of the following is the most basic oxide?(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the following orders is not in accordance with the property stated against it?(a) $F_2 > Cl_2 > Br_2 > I_2$ ; bond dissociation energy(b) $F_2 > Cl_2 > Br_2 > I_2$ ; oxidizing power(c) HI > HBr > HCl > HF; acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2$ ; electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) NF3(b) CIF3(c) BF3(d) AlF3Que 59. In which of the following pairs, the two species are isostructural?(a) BrO3 <sup>-</sup> and XeO3(b) SF4 and XeF4(c) SO3 <sup>2-</sup> and NO3 <sup>-</sup> (d) BF3 and NF3 <sup>-</sup>
(a) $PCl_5$ (b) $CCl_4$ (c) $PCl_3$ (d) $POCl_3$ Que 56. Which of the following is the most basic outlines to the most basic outlines to the most basic outlines to the most basic outlines.(a) $SeO_2$ (b) $Al_2O_3$ (c) $Sb_2O_3$ (d) $Bi_2O_3$ Que 57. Which of the following orders is not in accordance with the property stated against it?(a) $F_2 > Cl_2 > Br_2 > I_2;$ boord dissociation energy(b) $F_2 > Cl_2 > Br_2 > I_2;$ oxidizing power(c) $HI > HBr > HCl > HF;$ actic property power(d) $F_2 > Cl_2 > Br_2 > I_2;$ electronegativityQue 58. In which of the following molecules, are albords not equal?(a) $NF_3$ (b) $ClF_3$ (c) $BF_3$ (d) $AlF_3$ (b) $ClF_3$ (c) $BF_3$ (d) $Br_3$ and $NF_3^-$ (b) $SF_4$ and $XeF_4$ (c) $SO_3^{2^-}$ and $NO_3^-$
Que 56. Which of the following is the most basic oxide?(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the following orders is not in accordance with the property stated against it?(a) $F_2 > Cl_2 > Br_2 > I_2$ ; bond dissociation energy(b) $F_2 > Cl_2 > Br_2 > I_2$ ; oxidizing power(c) HI > HBr > HCl > HF; acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2$ ; electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) NF3(b) ClF3(c) BF3(d) AlF3Que 59. In which of the following pairs, the two species are isostructural?(a) BrO3 <sup>-</sup> and XeO3(b) SF4 and XeF4(c) SO3 <sup>2-</sup> and NO3 <sup>-</sup> (d) BF3 and NF3 <sup>-</sup>
(a) SeO2(b) Al2O3(c) Sb2O3(d) Bi2O3Que 57. Which of the following orders is not in accordance with the property stated against it?(a) $F_2 > Cl_2 > Br_2 > I_2;$ bond dissociation energy(b) $F_2 > Cl_2 > Br_2 > I_2;$ oxidizing power(c) HI > HBr > HCl > HF; acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2;$ electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) NF3(b) ClF3(c) BF3(d) AlF3Que 59. In which of the following pairs, the two spectrum s
Que 57. Which of the following orders is not in accordance with the property stated against it?(a) $F_2 > Cl_2 > Br_2 > I_2$ ; bond dissociation energy(b) $F_2 > Cl_2 > Br_2 > I_2$ ; oxidizing power(c) HI > HBr > HCl > HF; acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2$ ; electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) NF_3(b) CIF_3(c) BF_3(d) AIF_3Que 59. In which of the following pairs, the two species are isostructural?(a) BrO_3 <sup>-</sup> and XeO_3(b) SF_4 and XeF_4(c) SO_3 <sup>2-</sup> and NO_3 <sup>-</sup> (d) BF_3 and NF_3 <sup>-</sup>
(a) $F_2 > Cl_2 > Br_2 > I_2$ ; bond dissociation energy(b) $F_2 > Cl_2 > Br_2 > I_2$ ; oxidizing power(c) $HI > HBr > HCl > HF$ ; acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2$ ; electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) $NF_3$ (b) $ClF_3$ (c) $BF_3$ (d) $AlF_3$ Que 59. In which of the following pairs, the two species are isostructural?(a) $BrO_3^-$ and $XeO_3$ (b) $SF_4$ and $XeF_4$ (c) $SO_3^{2-}$ and $NO_3^-$ (d) $BF_3$ and $NF_3^-$
(c) $HI > HBr > HCl > HF;$ acidic property power(d) $F_2 > Cl_2 > Br_2 > I_2;$ electronegativityQue 58. In which of the following molecules, are all the bonds not equal?(a) $NF_3$ (b) $ClF_3$ (c) $BF_3$ (d) $AlF_3$ Que 59. In which of the following pairs, the two species are isostructural?(a) $BrO_3^-$ and $XeO_3$ (b) $SF_4$ and $XeF_4$ (c) $SO_3^{2-}$ and $NO_3^-$ (d) $BF_3$ and $NF_3^-$
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(a) $BrO_3^-$ and $XeO_3$ (b) $SF_4$ and $XeF_4$ (c) $SO_3^{2-}$ and $NO_3^-$ (d) $BF_3$ and $NF_3^-$
Que 60. Which of the following increasing order is not correct as mentioned in the property with it?
(a) $HClO_2 < HClO_3 < HClO_4$ (thermal stability) (b) $HClO_4 < HClO_3 < HClO_2 < HClO$ (oxidising power)
(c) $F^- < Cl^- < Br^- < 1^-$ (reducing nature) (d) $HlO_4 < lCl < l_2 < Hl$ (oxidation number of iodine)